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8. Examine the orbital diagrams and electron configurations in Model 2. Using the following electron configuration:  $1s^2 2s^2 2p^4$  a. Underline the energy levels. b. Circle the sublevels. c. Draw a box around the numbers of electrons. 9. The 2s and 2p sublevels are very close in energy, as are the 3s and 3p sublevels. Explain how the

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Electron Configuration 5 11. The Pauli exclusion principle describes the restriction on the placement of electrons into the same orbital. The Pauli exclusion principle can be expressed as: "If two electrons occupy the same orbital, they must have \_\_\_\_." the same spin opposite spins 12. Hund's rule describes how electrons are distributed among orbitals of the same sublevel when there is ...

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The subshells have a distinct shape and configuration, in which the electrons move freely. Each shell and subshell have a limitation on the amount of electrons that it can carry. The maximum electrons that can be carried by the sub-shell S is 2, by P is 6, by D is 10, and the F sub-shell can carry 14. This decides the electron capacity of the ...

### **Electron Configuration Chart for All Elements in the ...**

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