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STRUCTURE OF ATOM 27 ••• to explain the for mation of dif fer ent kinds of molecules by the combination of different atoms and, ••• to underst and the origin and natur e of the characteristics of electromagnetic radiation absorbed or emitted by atoms.
2.1 SUB-ATOMIC PARTICLES Dalton's atomic theory was able to explain

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1. The charge on an electron is _____. (-2.46 x 10⁴ coulombs, -1.6 x 10⁻¹⁹ coulombs, 1.6 x 10⁻⁹ coulombs)
2. The maximum number of electrons that can accommodated by a p-orbital is _____.

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Which one is the wrong statement? (a) de-Broglie's wavelength is given by $\lambda = h/mv$. where m = mass of the particle, v = group velocity of the particle (b) The uncertainty principle is $\Delta E \times \Delta t \geq h / 4 \pi$ (c) Half-filled and fully filled orbitals have greater stability due to greater exchange energy, greater symmetry and more balanced arrangement

NEET and AIPMT NEET Chemistry Structure of Atom MCQ ...

CHEM1001 June 2008 Example Multiple Choice Questions 1. Which one of the following best defines the word "allotropes"? A Different structural forms of an element B A pair of substances that differ by H^+ C Elements that possess properties intermediate between those of metals and non-metals D Atoms of a given atomic number that have a specific number of neutrons

The following multiple choice questions are provided to ...

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Chemistry Questions And Answers On Atomic Structure

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Composition of atom Electron ($-1e_0$) (1)

It was discovered by J.J. Thomson (1897) and is negatively charged particle. (2)

Electron is a component particle of cathode rays.

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rutherford model of atom, atomic emission

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Chemistry MCQs For Entry Test Preparation (Chapter Wise)

(g) Atoms of same element with same atomic number but a different mass number. (h) The smallest unit of an element which takes part in a chemical reaction. Answer: (a) Neutron (b) Mass number (c) Valency (d) Orbits or Valence shells (e) Atomic number (f) Mass number (g) Isotopes (h) Atom. 4.
Multiple Choice Questions

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